

Dual functional self-regenerative Cr-substituted Ba₂In₂O₅

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Ba₂In₂O₅ is a high-temperature oxide and proton conductor with mixed ionic-electronic conductivity[1]. In such materials ceramic membranes for oxygen gas separation is one of the interesting research areas[2]. The yellow colored powders suggest that it is potentially applicable as a photocatalyst for the activation of small molecules such as H₂O and CO₂[3,4]. The partial substitution of In³⁺ by Cr³⁺ altered the crystal structure and changed the optical bandgap. The crystal structure changes are important for membrane applications, while the resulting bandgap variations are relevant for photocatalyst applications.

Addition of Cr³⁺ induced a crystal structure change from *Ibm2* to *I4cm* due to an enhanced disorder of oxygen vacancies. This increased disorder resulted in a reduced activation energy for oxygen transport *via* oxygen vacancies. However, more than 10 % Cr substitution for In caused a shrinkage of the unit cell resulting in a reduced self-diffusion coefficient of the oxide ions. Consequently, Ba₂In_{1.8}Cr_{0.2}O₅ exhibited the highest oxygen permeability with $P(\text{O}_2) = 1.4 \text{ mL} \cdot \text{cm}^{-1} \cdot \text{min}^{-1} \cdot \text{mm}$ at $T = 1223 \text{ K}$ (Fig. 1). Thermochemical studies showed an easy membrane material reformation using the end-of-lifetime phases also suggesting a potential self-regeneration behavior using a regular oxygen treatment[5].

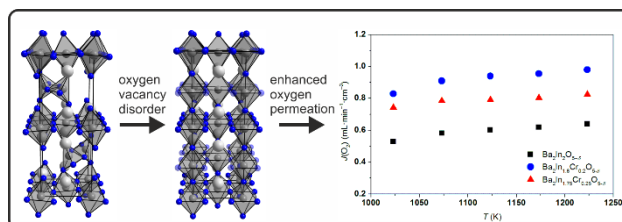


Figure 1: Crystal structure change / oxygen vacancy disorder (left) and oxygen permeability (right) [2].

Furthermore, increasing Cr³⁺ substitution reduced the bandgap and improved the photocatalytic conversion of CO₂ in presence of H₂. Ba₂In_{1.4}Cr_{0.6}O₅ exhibited a compatible photoactivity with that of reference P-25 (TiO₂) (Fig. 2), even though having a 250-times smaller BET surface area[4].

Even though Cr-substituted Ba₂In₂O₅ involves the usage of the critical element indium, it can be interesting to be used, since it allows high oxygen permeation at ca. 200 K lower T than classical membrane materials [2]. Additionally, the indium can be easily recovered in the material cycle, e.g. the recyclate of used membranes can be again applied as photocatalyst.

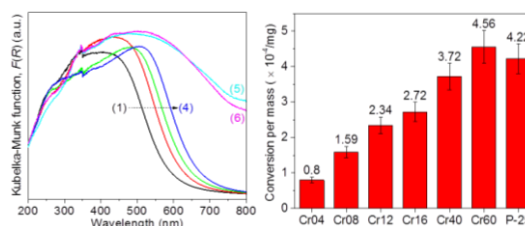


Figure 2: Optical bandgaps of Ba₂In_{2-x}Cr_xO₅ (left; $x = 0.04$ (1), 0.08 (2), 0.12 (3), 0.16 (4), 0.4 (5), 0.6 (6)) and photocatalytic CO₂ conversion compared with P-25 as reference material (right) [4].

[1] J. Goodenough et al., *Solid State Ionics* **52** (1992), 105.

[2] M. Widenmeyer et al., *J. Mem. Sci.* **595** (2020), 117558.

[3] S. Yoon et al., *Solid State Sci.* **73** (2017), 1.

[4] S. Yoon et al., *Solid State Sci.* **78** (2018), 22.

[5] M. Widenmeyer et al., *Thermochim. Acta* (2019) submitted.

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